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Pogil Colligative  
Properties  
Answers

# **Pogil Colligative Properties Answers**

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Colligative Properties –  
Part 1 Why? There is a  
general misconception  
that adding sodium  
chloride to cooking  
water for pasta  
increases the  
temperature of the  
boiling water so that it  
cooks the pasta faster.  
Adding sodium chloride  
to boiling water does  
increase the

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**Colligative**

**Properties Part 1**

Model 2: Ionic Solutes  
and Colligative

Properties A colligative  
property is a property  
that is affected by the  
concentration of solute  
particles dissolved in  
solution. Key Question

6. Explain why the  
freezing point of

0.0302 m KI (aq) is

-0.111°C, but the

freezing point for the

0.0302 m aqueous

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sucrose solution is  
-0.056°C. PRACTICE B.

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**POGIL: Colligative**  
**Properties Part 2**

called colligative properties. Examples of colligative properties include vapor pressure, boiling point, freezing point, and osmotic pressure. For discussions of colligative properties it is often convenient to express solution

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concentration in terms  
of molality; The  
magnitude of any  
colligative effect  
depends on the total  
concentration of ...

**Chem 116 POGIL  
Worksheet - Week 4  
Properties of  
Solutions**

Question: POGIL:  
Colligative Properties -  
Part 1 | Page 6  
PRACTICE B. Ethylene  
Glycol (HOCH<sub>2</sub>CH<sub>2</sub>OH)  
Is A Molecular



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Substance That Is  
Added To That Water  
Put In An Automobile's  
Radiator. One Of The  
Advantages  
"antifreeze" Offers Is  
That The Water Will  
Boil At Higher  
Temperature Than Its  
Normal Boiling Point,  
Helping To Prevent  
Engine Overheating.

**Solved: POGIL:  
Colligative  
Properties - Part 1 |  
Page 6 PR ...**

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A walkthrough of the responses for the POGIL on colligative properties that we worked together in small groups in class this week.

## **POGIL Colligative Properties Overview**

WORKSHEET: SOLUTIONS AND COLLIGATIVE PROPERTIES SET A: 1.

Find the molarity of all ions in a solution that contains 0.165 moles of aluminum chloride in

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820. ml solution.

Answer:  $[Al^{3+}] = 0.201$

M,  $[Cl^-] = 0.603$ M. 2.

Find the molarity of each ion present after mixing 27 ml of 0.25 M  $HNO_3$  with 36 ml of 0.42 M  $Ca(NO_3)_2$

**WORKSHEET: SOLUTIONS AND COLLIGATIVE PROPERTIES SET A**

6. Looking at your answers to Question 1 and 2 from Model 1, tell what atoms are

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represented by: a. The small, unshaded circles in Model 2. The small unshaded circles represent hydrogen. b. The larger gray shaded circles in Model 2. The larger gray shaded circles represent oxygen. 7. What do the solid lines between the small and large circles ...

## **Properties of Water**

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properties answers is

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## **POGIL: Le Chatelier's Principle Why? MODEL #1: Disrupting ...**

POGIL: There's Power  
in Numbers Honors

Chemistry Unit 11:

Solutions Page 1 of 4

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## Pogil Colligative Properties

Name: Group

Members: ! Learning

Objective: □ Identify the basic colligative properties. □ Calculate the effects of solutes on boiling point, freezing point, and osmotic pressure.

### **Honors Chemistry**

### **POGIL: There's**

### **Power in Numbers**

Which of the following colligative properties would a sand and water mixture exhibit?

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- a. a lowering of the water's freezing point  
b. an osmotic pressure  
c. both of the above are correct  
d. none of the above are correct

**Answered: Which of the following colligative... | bartleby**

NAME AP Chemistry  
POGIL Colligative  
Properties Part 2 April  
27th, 2019 - Model 2  
Ionic Solutes and  
Colligative Properties A

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colligative property is a property that is affected by the concentration of solute particles dissolved in solution Key Question 6 Explain why the freezing point of 0.0302 m KI aq is  $0.111^{\circ}\text{C}$  but the freezing point for the 0.0302 m aqueous sucrose solution is  $0.056^{\circ}\text{C}$  ...

**Molality and colligative properties answer**

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## key

Play this game to review Quantitative Chemistry. What observation(s) reflect(s) colligative properties? I. A 0.5 m NaBr soln has a higher vapor pressure than 0.5 m BaCl<sub>2</sub> soln. II. A 0.5 m NaOH soln freezes at a lower T. than pure water. III. Pure water freezes at a higher T than pure methanol.

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**Colligative  
Properties Quiz -  
Quizizz**

Using colligative properties to calculate the molar mass of a nonvolatile, non-electrolyte. One of the most important applications of colligative properties is that they can be used to determine molar mass. This is done as follows: A known mass of a substance is dissolved in a known

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volume of solution or  
mass of solvent.

**Colligative  
Properties  
(Worksheet) -  
Chemistry  
LibreTexts**

Chem 116 POGIL  
Worksheet - Week 3  
Intermolecular Forces,  
Liquids, Solids, and  
Solutions ... †  
Understand the effect  
of solute concentration  
on colligative  
properties ... † Be able



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to calculate the  
magnitude of  
colligative changes in  
vapor pressure,  
freezing point, boiling  
point, and osmotic  
pressure

## **Chem 116 POGIL Worksheet - Week 3 Intermolecular Forces ...**

To get started finding  
Physical And Chemical  
Properties Pogil  
Answers , you are right  
to find our website

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