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□□stoichiometry The calculation of quantities in chemical reactions. atoms, molecules, moles, mass, and volume  
The quantities to consider in a chemical  
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25 terms. jakeydv. chapter 12 study guide chemistry. law of conservation of mass. formula mass. law of definite

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proportions. limiting reactant. no detectable change in mass occurs during a chemical reaction. the sum of the atomic masses given in the formula of a substan....

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Chapter 12 - States of Matter Section  
12.1 Gases 1. motion 2. a. small b.  
forces c. random d. elastic; kinetic 3.  $KE = \frac{1}{2}mv^2$  4. Temperature 5. true 6. true  
7. false 8. true 9. true 10. false 11. true  
12. false 13. a 14. a 15. d 16. d 17. b 18.  
b ...

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CHAPTER Section 12.4 Phase Changes In  
your textbook, read about phase  
changes. Date Class STUDY GUIDE  
Complete the table by writing the initial  
and final phases for each phase change  
and making a check (V) in the correct  
energy column. Phase Change Phase  
Energy released required initial final SO  
Column B boiling point freezing point  
melting point

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## **Chapter 12 - Chemical Kinetics | CourseNotes**

1. Chapter 12 - Chemical Kinetics. 12.1 Reaction Rates. A. Chemical kinetics 1. Study of the speed with which reactants are converted to products B. Reaction Rate 1. The change in concentration of a reactant or product per unit of time.  $t$  A  $t$  concentration of A at time  $t$

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concentration of A at time t Rate. 2 1 2  
1.

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Matter and Change • Chapter 12 Section 12.2 Stoichiometric Calculations In your textbook, read about mole-to-mole conversion. Read the following passage and then solve the problems. In the equation that follows each problem, write in the space provided the mole ratio that can be used to solve the problem.

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