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$K_3 [CrI_6] [Cu(en)_2 (H_2O)_2]Cl_2; Na_3$

$[Co(NO_2)_6]$  Solution.

The complexes are octahedral.  $Cr^{3+}$  has a  $d^3$  configuration.

These electrons will all be unpaired.  $Cu^{2+}$  is  $d^9$ , so there will be one unpaired electron.  $Co^{3+}$  has  $d^6$  valence electrons, so the

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Crystal field splitting  
will determine how  
many are paired.

Nitrite is a strong-field

...

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(continued) 15. Write the chemical symbol for the ion with 21.

How many protons, neutrons, and 29 protons and 27 electrons, electrons are present in the  $n_4$  ion?

16. How many protons, neutrons, and 22. Write the complete chemical symbol for electrons are present in the  $Ni_2$  ion? the ion with 84 protons, 125 neutrons, and

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**Name Date Class**

**3—3 Practice**

**Problems - Weebly**

These are homework  
exercises to

accompany the

Textmap created for

"Chemistry: The

Central Science" by

Brown et al.

Complementary

General Chemistry

question banks can be

found for other

Textmaps and can be

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In addition to these publicly available questions, access to private problems bank for use in exams and homework is available to faculty only on an individual basis; please ...

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This is a collection of worked general chemistry and

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introductory chemistry problems, listed in alphabetical order. Included are printable pdf chemistry worksheets so you can practice problems and then check your answers. You may also browse chemistry problems according to the type of problem.

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Organic Chemistry

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Practice Problems. The problem sets provided here are similar to those found on various kinds of standardized exams, such as GRE, ACS & MCAT. The questions are roughly organized by subject, and most sets have over 50 multiple choice problems. To use a problem set, click on its descriptive title.

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## Varsity Tutors

$V = 15.24 \text{ cm} \times 10.16 \text{ cm} \times 5.08 \text{ cm}$   
 $V = 786.58 \text{ cm}^3$  Place this into the formula above:  
 $m = DV$   $m = 19.3 \text{ g/cm}^3 \times 786.58 \text{ cm}^3$   
 $m = 14833.59 \text{ grams}$  The answer we want is the mass of the gold bar in kilograms. ... For more example problems, use the Worked Chemistry Problems. It contains over 100 different worked example problems useful for

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**Calculate Mass from  
Density Example  
Problem**

Name Date 19—1

Practice Problems 1. 2.

3. 4. 5. 8 What is the

concentration of OH-

ions in saturated

limewater if  $[H^+] = 3.98$

$\times 10^{-11}$  M? Is acidic, basic, or

neutral?

**Kennedy HS**

Example #13: A golden-

colored cube is handed

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to you. The person wants you to buy it for \$100, saying that is a gold nugget. You pull out your chemistry text and look up gold's density and read that its density is  $19.32 \text{ g/cm}^3$ . You measure the cube and find that it is  $2.00 \text{ cm}$  on each side, and weighs  $40.0 \text{ g}$ .

### **ChemTeam: Density**

Organic Chemistry  
Practice Problems at

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Michigan State  
University. The

following problems are meant to be useful study tools for students involved in most undergraduate organic chemistry courses. The problems have been color-coded to indicate whether they are: 1. Generally useful, 2.

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Edition offers the full  
text plus the full  
solutions manual and  
study guide integrated  
throughout. The  
solutions are authored  
by David Klein and  
very robust to provide  
students with the  
guidance and support  
they need to be  
successful in this  
course. Success in  
organic chemistry

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Chemistry  
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requires mastery in  
two core aspects ...

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Practice Problem 5.

What is the value of a gold ingot 20.0 cm long by 8.5 cm wide by 6.0 cm tall, if the mass of a cubic centimeter of gold is 19.3 grams and the price of gold is \$356 per ounce? Click here to check your

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answer to Practice  
Problem 5

## **Review of Measurement - Purdue University**

HW2 Practice Problem  
WS: 5-10 HW3 Practice  
Problem WS: 1-4, 11  
HW4 "10-2 Practice  
Problems":

1,3,6,8,18,19 pg. 191:  
33,34; pg. 194: 38

HW5 "Formula  
Stoichiometry WS"  
1-19 HW6 "Ch 8  
Practice Problems WS"

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**Bader, Mr. K. -  
Science / Honors  
Chem Homework**

Chemistry 13 4  
Practice Problems  
Answer Key solute is  
volatile, which usually  
describes miscible  
liquids, you can have  
all mole fraction ranges  
in Figure 13.4.4, but  
both substances  
contribute to the total  
vapor pressure. 13.4:  
Colligative Properties -



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